# Phase Diagram

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## Binary Phase Diagram



- Constituent- Components in the solution
- Composition-Any ratio of components
- Composition (40 wt.% Ni+60 wt% Cu) for example

# Binary Phase Diagram with Equilibrium Cooling



- At 35 wt% Ni-65 wt% Cu, cooling from 1300 C
- Consider Point a, b,c,d,e

### Concept Check

A copper-nickel alloy of composition 70 wt% Ni-30 wt% Cu is slowly heated from a temperature of 1300°C (2370°F).

- (a) At what temperature does the first liquid phase form?
- (b) What is the composition of this liquid phase?
- (c) At what temperature does complete melting of the alloy occur?
- (d) What is the composition of the last solid remaining prior to complete melting?



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- a. 1350 C
- b. WL= 55 wt% Ni, 45 wt% Cu
- c. 1380 C
- d. Ws= 80 wt% Ni, 20 wt% Cu



#### Determination of Phase Amount

- Consider at 1250 C (Point B) where Alpha+ Liquid Phases present
  - Determine W<sub>L</sub> and W<sub>alpha</sub> at Point B
  - Composition need to determine in terms of only one constituent

• Lever Rule





W<sub>L</sub>=1-Ws



Alpha Phase is solid Pahse

At Point B, Ni wt.%, C<sub>0</sub>=35 wt.% C<sub>L</sub>=32 wt.% C <sub>alpha</sub>=43%

H.W.  $\rightarrow$  Determine Solid and Liquid Phases at 1225 C

 $W_L = \frac{42.5 - 35}{42.5 - 31.5} = 0.68 \text{ or } 68 \text{ wt.\%}$  Liquid, so reminder 32 wt% Solid alpha phase



- Polymorphic Transformation-
- Consider Pure Iron
  - Ferrite- A stable iron phase, Alpha Iron.
  - At 912 C-- Ferrite to Austenite, FCC gamma iron
  - At 1394 Austenite to Delta Ferrite, BCC

HW→Now if you consider Carbon in it. For example, Determine Phase Changes at 0.5 wt% C As you heat up from Room to 1600 C. See what happens

#### Homework



Fig-Partial Cu-Ni Phase Diagram

- Q1. In the Copper-Nickel Phase Diagram, Determine Solid and Liquid Phase Amount of composition B at 1225 C.
- Q2. From Iron-Carbon Phase Diagram, Determine Phase Changes at 0.5 wt% C As you heat up from Room to 1600 C.